

## Connect the Dots HW

Read and outline (Cornell Style) **Section 9.1** in your textbook. Then answer the following assessment questions:

1. What is a covalent bond? How does it differ from an ionic bond?
2. What is a single covalent bond? Why does it form?
3. Why do multiple bonds form?
4. Draw the Lewis Dot Diagrams for the following elements AND determine how many covalent bonds each would make:
  - a. Se
  - b. Br
  - c. K
  - d. As
  - e. Ga
  - f. Ge
5. Draw the Lewis Dot Diagrams for the following molecules:
  - a. HBr
  - b. SbH<sub>3</sub>
  - c. H<sub>2</sub>Se
  - d. SiCl<sub>2</sub>H<sub>2</sub>
  - e. CH<sub>3</sub>NH<sub>2</sub>
  - f. HOCl
  - g. CH<sub>2</sub>Cl<sub>2</sub>